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ELEMENT TO MASS Sample Problem 4 uses the progression of steps 1→2→3 to convert from the mass of an element to the number of atoms. In order to calculate the mass from a given number of atoms, these steps will be reversed.

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Only 0.0121 mol Cl₂ is available. For all
of the KBr to react, 0.0136 mol Cl₂ is
needed. Therefore, Cl₂ is the limiting
reactant. EVALUATE Is the determination
of limiting reactant reasonable? Yes; the
mass of 2 mol KBr is 238 g and the mass
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by reacting nitrogen and hydrogen
under pressure, at high temperature,

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and in the presence of a catalyst. The equation is $N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$. If 4.0 mol of H_2 react, how many moles of NH_3 will be produced?

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